UNIT-2

ELECTROCHEMISTRY AND CORROSION

Short Answer Questions:

1. Define Electrode potential.
2. Write a Nernst equation.
3. What is Standard electrode potential.
4. Define Electrochemical cell.
5. Write any two applications of salt bridge.
6. What is calomel electrode.
7. Define dry theory of corrosion.
8. Write any two effects of corrosion.
9. Write any two factors that affect rate of corrosion.
10. What is Standard Hydrogen Electrode

Long answers question:

1. What is Electrochemical cell.Write construction and working of a Galvanic cell.
2. Write the construction, working of a Standard Hydrogen Electrode.
3. Write the construction, working of a Calomel electrode.
4. Explain Chemical theory of a Corrosion.
5. Explain Electrochemical theory of a Corrosion.
6. Explain different types of Corrosion.
7. What are the factors which influence the Rate of Corrosion.
8. What is meant by cathodic protection. Explain sacrificial and impressed cathodic protection.
9. Define corrosion. Write the causes and effects of corrosion.
10. The standard reduction potential of the zinc electrode and copper electrode -0.76V and +0.34V respectively. Find the standard emf of the cell.